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Laboratory manual for physical chemistry where  $T_0$  is the freezing point of the solvent. If BC in Fig. 1st-1 is taken to be a straight line, the slope  $dp_A/dT$  may be replaced by  $CB/BD$ , i.e., by  $(p_A - p_s)/\Delta T$ , so that  $p_p = p_s + \Delta T \cdot \frac{dp_A}{dT}$  Eq. 1st.4 The slope of the vapor pressure curve of the solid

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solvent is also given by another form of the

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The dilution formula allows you to determine the molarity of a solution after it has been diluted. The formula is:  $M_1 V_1 = M_2 V_2$  Where  $M_1$  and  $V_1$  are the molarity and volume before a dilution, and  $M_2$  and  $V_2$  are the molarity and volume after a dilution. Consider this example from the 2006 NECTA:  
"Hydrochloric acid solution made by diluting

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1.1 Packing models of (a) side-by-side layer, (b) hexagon  
layer, (c) simple cubic layer, (d) body centered cubic.  
Determine the maximum amount of balls that can touch  
another one ball (center- ball) in the same layer for each  
arrangement.

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